

CBCS SCHEME

21CHE12/22



First/Second Semester B.E. Degree Examination, June/July 2023 Engineering Chemistry

Time: 3 hrs.

Max. Marks: 100

Note: Answer any FIVE full questions, choosing ONE full question from each module.

Module-1

- Define single electrode potential. Derive Nernst equation for single electrode potential. (07 Marks)
 - What are ion selective electrodes? Explain the determination of pH using Glass Electrode. (07 Marks)
 - Distinguish between primary, secondary and reserve batteries. (06 Marks)

OR

- Describe the construction and working of Li-ion battery. Mention its applications. (07 Marks)
 - What voltage will be generated by a cell that consists of iron electrode immersed in 0.1M FeSO_4 solution and a silver electrode immersed in 0.05M AgNO_3 solution at 298K. given standard electrode potentials of Fe and Ag are -0.44V and 0.80V respectively. Write the cell representation and cell reactions. (07 Marks)
 - What are reference electrodes? Explain the construction and working of calomel electrode. (06 Marks)

Module-2

- What is corrosion? Describe the electrochemical theory of corrosion by taking iron metal as an example. (07 Marks)
 - Explain the factors affecting the rate of corrosion :
 - Nature of corrosion product
 - Ratio of anode to cathodic areas
 - pH. (07 Marks)
 - What is electroless plating? Outline the electroless plating of copper. (06 Marks)

OR

- What is meant by metal finishing? Mention (any five) technological importance of metal finishing. (06 Marks)
 - What is electroplating? Discuss the electroplating of chromium. (07 Marks)
 - Explain the process of :
 - Galvanizing process
 - Anodizing of Aluminium. (07 Marks)

Module-3

- What are polymer composites? Explain the synthesis and application of Kevlar fibre. (06 Marks)
 - What are conducting polymers? Describe the mechanism of conduction in poly Aniline. (07 Marks)
 - Briefly explain the carbon nanotubes with properties and applications. (07 Marks)

Important Note : 1. On completing your answers, compulsorily draw diagonal cross lines on the remaining blank pages.
2. Any revealing of identification, appeal to evaluator and /or equations written eg, $42+8=50$, will be treated as malpractice.

OR

- 6 a. Describe the synthesis of nano-material by sol-gel technique. (07 Marks)
b. Explain any three size dependent properties of nano material. (06 Marks)
c. Explain the synthesis, properties and application of polyurethane. (07 Marks)

Module-4

- 7 a. Briefly explain any six basic principles of green chemistry. (06 Marks)
b. Explain the synthesis of paracetamol by conventional and green route from phenol. (07 Marks)
c. What are PV cells? Describe the construction and working of photovoltaic cells. (07 Marks)

OR

- 8 a. With a neat diagram, explain the production of hydrogen by photocatalytic method. (07 Marks)
b. Explain the following with example :
i) Solvent free reaction (07 Marks)
ii) Micro wave synthesis. (06 Marks)
c. Describe the construction and working of methanol-oxygen fuel cell. (06 Marks)

Module-5

- 9 a. Explain the theory, instrumentation and application of colorimetry. (07 Marks)
b. Explain the determination of hardness of water by EDTA method. (07 Marks)
c. In c COD test 28.5cm^3 and 13.5cm^3 of 0.05N FAS solutions are required for blank and sample titration respectively. The volume of test sample used is 25cm^3 . Calculate the COD of the sample solution. (06 Marks)

OR

- 10 a. Define the following units of standard solution :
i) Normality (06 Marks)
ii) Molarity (07 Marks)
iii) PPM. (07 Marks)
b. Define COD. Explain the determination of COD of waste water sample. (07 Marks)
c. Explain the theory, instrumentation and application of flame photometry. (07 Marks)
